

Centre Number	Candidate Number	Name
---------------	------------------	------

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
General Certificate of Education Ordinary Level

**CHEMISTRY**

**5070/02**

Paper 2

October/November 2004

**1 hour 30 minutes**

Candidates answer on the Question Paper.  
Additional Materials: Answer Paper.

**READ THESE INSTRUCTIONS FIRST**

Write your name, Centre number and candidate number in the spaces provided at the top of this page and on any separate answer paper used.

Write in dark blue or black pen in the spaces provided on the Question Paper.

You may use a pencil for any diagrams, graphs, or rough working.

You may use a calculator.

Do not use staples, paper clips, highlighters, glue or correction fluid.

**Section A**

Answer **all** questions.

A copy of the Periodic Table is printed on page 16.

**Section B**

Answer **three** questions.

Write your answers on any line pages provided and/or a separate answer paper.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

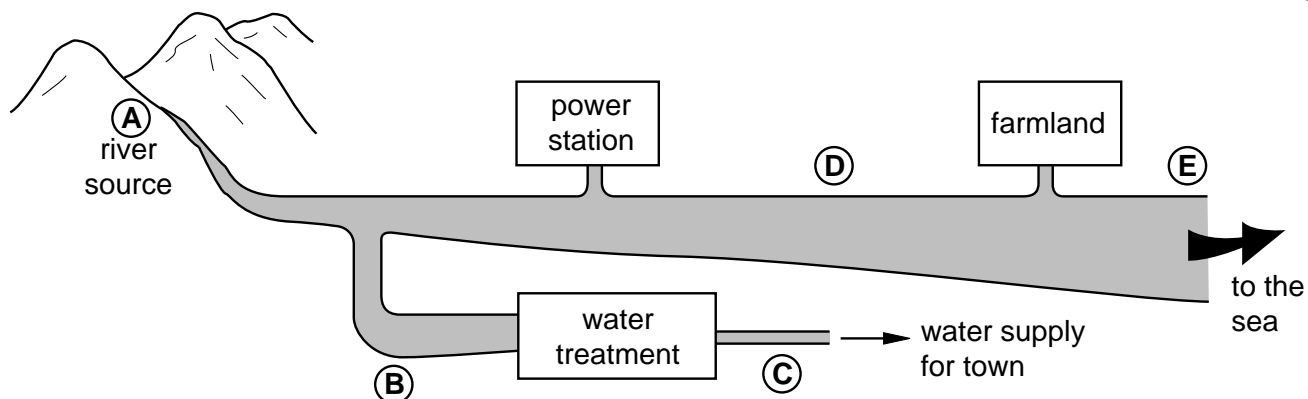
For Examiner's Use	
<b>Section A</b>	
<b>B7</b>	
<b>B8</b>	
<b>B9</b>	
<b>B10</b>	
<b>TOTAL</b>	

If you have been given a label, look at the details. If any details are incorrect or missing, please fill in your correct details in the space given at the top of this page.

Stick your personal label here, if provided.

Section A

A1 The diagram shows where five water samples, A to E, were taken from a river.



The table shows information about the water samples.

sample	temperature / °C	dissolved oxygen / ppm
A	6	15
B	5	13
C	6	13
D	13	12
E	8	

(a) Describe how the temperature of the river water changes as it flows from the source of the river to the sea.

..... [1]

(b) Fertiliser enters the river as it flows past the farmland.

(i) Suggest the oxygen content of water sample E.

.....

(ii) Explain your reasoning.

.....

..... [3]

- (c) Samples **B** was taken before and sample **C** was taken after the water was treated. The water is then used as the water supply for the town. Complete the table to show how the contents of the water change when the water is treated.

<i>contents</i>	<i>change</i> <i>(increases / decreases / stays the same)</i>
dissolved minerals	<i>stays the same</i>
suspended particles	
dissolved oxygen	<i>stays the same</i>
living microbes (e.g. bacteria)	
chlorine	

[3]

A2

<i>substance</i>	<i>type of bonding</i>	<i>melting point / °C</i>	<i>boiling point / °C</i>
iodine	covalent	114	184
lead(II) bromide	ionic	370	914
methane	covalent	-182	-161
bromine	covalent	-7	59
silicon dioxide	covalent	1610	2230
lithium	metallic	180	1360

Use the substances named in the table to answer the following questions.

(a) Name the substances that are **not** solids at room temperature and pressure.

..... [1]

(b) Which substance is a liquid over the largest temperature range?

..... [1]

(c) Name the substances that are non-metallic elements.

..... [1]

(d) Which **two** substances conduct electricity when molten?

..... [1]

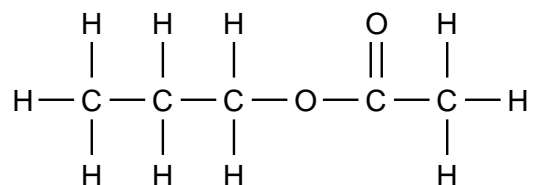
(e) Explain, using ideas about structure, why methane and silicon dioxide have different melting points.

.....  
 ..... [2]

(f) Describe a method for making lead from lead(II) bromide.

.....  
 ..... [2]

- A3** This is the structure of an ester made in a reversible reaction between a carboxylic acid and an alcohol.



- (a) (i)** State the conditions for this reaction.

.....  
 ..... [2]

- (ii)** Draw the structure of the carboxylic acid used in the reaction.

[1]

- (iii)** Write an equation for this reaction.

..... [2]

- (b)** A student carried out some experiments to compare the relative strengths of dilute ethanoic acid with dilute hydrochloric acid.

- (i)** Describe a test that can be used to distinguish between dilute ethanoic acid and dilute hydrochloric acid.

.....  
 ..... [2]

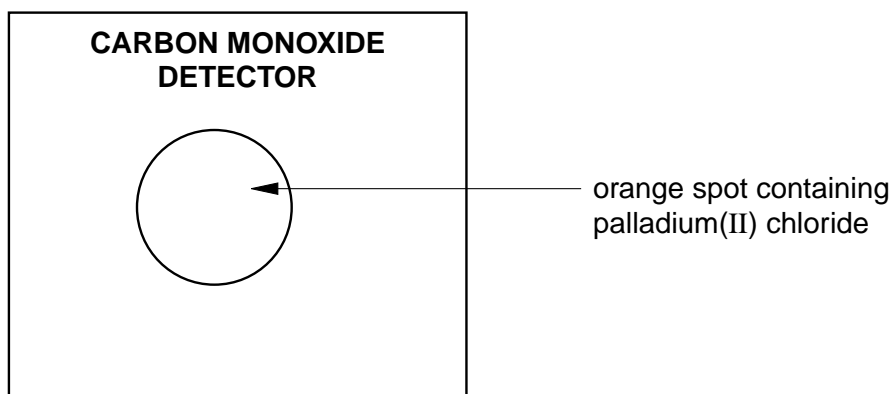
- (ii)** Name a solid substance that will react with both acids. Describe what you will **see** during the reaction.

substance .....

observations .....

..... [2]

A4 Carbon monoxide detectors can be used in the home.

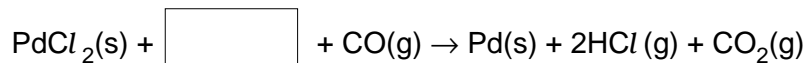


The orange spot turns black if there is a high concentration of carbon monoxide in the air.

(a) Why is carbon monoxide hazardous?

..... [1]

(b) The spot turns black when palladium(II) chloride reacts with carbon monoxide to form palladium metal.



(i) Complete the equation by writing the formula of the missing reactant in the box.

(ii) Complete the table to show the oxidation states of palladium and carbon before and after the reaction takes place.

<i>element</i>	<i>oxidation state before reaction</i>	<i>oxidation state after the reaction</i>
palladium		
carbon		

(iii) Use information from the table to explain why this is a redox reaction.

.....  
..... [5]

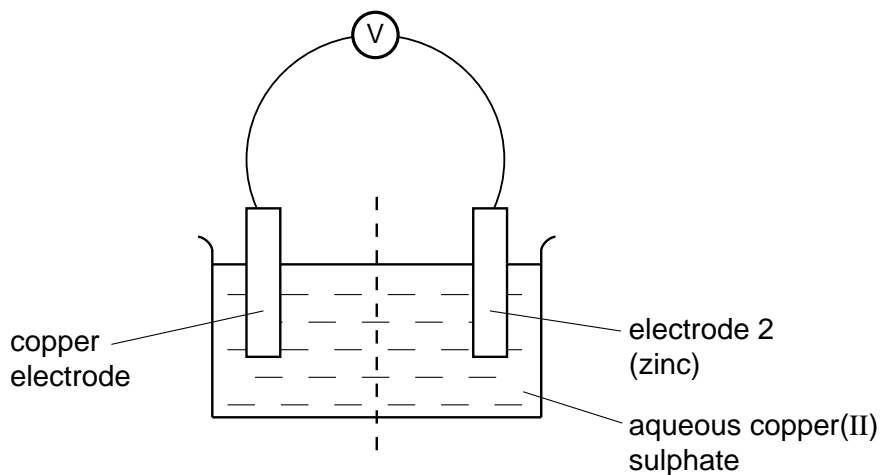
(c) Name **one** industrial process that uses carbon monoxide as a reducing agent.

..... [1]

A5 (a) Write an ionic equation for the reaction between zinc and aqueous copper(II) sulphate.

..... [1]

This reaction can be used to generate electricity in a cell.



- (b) Draw an arrow on the diagram to show the direction of the flow of electrons in the wire. [1]
- (c) The voltage of the cell was measured when the following metals were used as electrode 2.

**copper      iron      lead      zinc**

Complete the table by entering the metals in the correct order.

<i>meter reading / V</i>	<i>metal</i>
1.10	
0.78	
0.21	
0.00	

[2]

- (d) When **metal M** was used as electrode 2, it produced a higher voltage than zinc. Suggest a name for **metal M**.

..... [1]

**A6** This question is about making salts.

**(a)** For each salt, suggest the name of the missing reagent and briefly describe how obtain the solid product from the reaction mixture.

**(i)** Salt to be made: lithium chloride.

reagent 1: dilute hydrochloric acid

reagent 2: .....

I could obtain solid lithium chloride by: .....

.....

**(ii)** Salt to be made: barium sulphate.

reagent 1: aqueous potassium sulphate

reagent 2: .....

I could obtain solid barium sulphate by: .....

.....

**(iii)** Salt to be made: blue copper(II) sulphate crystals.

reagent 1: dilute sulphuric acid

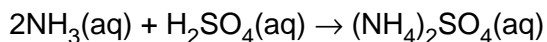
reagent 2: .....

I could obtain blue copper(II) sulphate crystals by:

.....

..... [6]

**(b)** Ammonium sulphate can be made by reacting aqueous ammonia with dilute sulphuric acid.



Calculate the mass of ammonium sulphate that can be made from 51 g ammonia.

.....

.....

..... [3]



## Section B

Answer **three** questions from this section.  
Tie the extra sheets used loosely to this booklet.

**B7** Magnesium carbonate decomposes when it is heated.

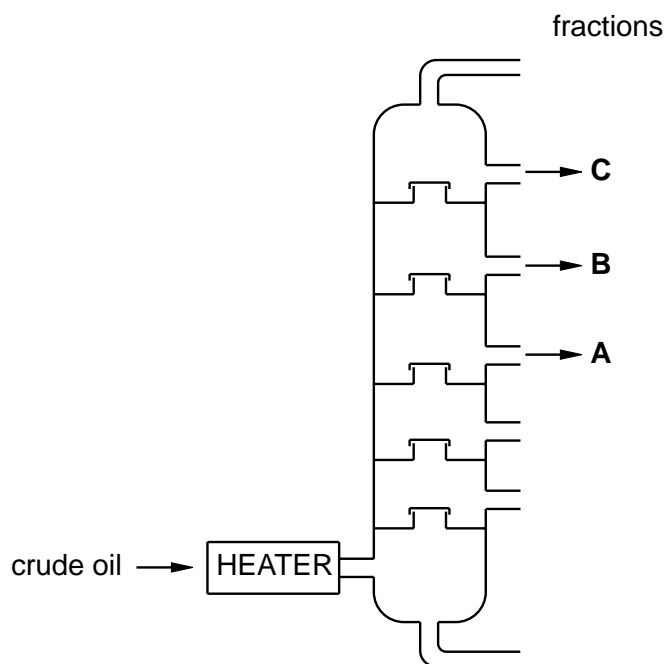


In an experiment, 10.5 g of magnesium carbonate was heated to a constant mass.

- (a) Sketch a graph to show how the volume of carbon dioxide collected changes with time. Explain your answer. [3]
- (b) Calculate the maximum volume of carbon dioxide, at room temperature and pressure, that can be formed from 10.5 g of magnesium carbonate. [3]
- (c) The experiment was repeated under the same conditions using zinc carbonate instead of magnesium carbonate.
- (i) Describe how the rates of the reactions would be different. Explain your answer.
- (ii) The same mass (10.5 g) of zinc carbonate was used. Would the total volume of carbon dioxide formed be the same? Explain your answer. [4]

[Total: 10 marks]

**B8** This diagram shows a fractionating column for the separation of crude oil.

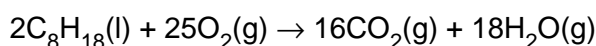


The following fractions leave the column.

<i>fraction</i>	<i>number of carbon atoms</i>	<i>boiling range / °C</i>
naptha	7 – 14	90 – 150
paraffin	9 – 16	150 – 240
diesel oil	15 – 25	220 – 250

- (a) Which fractions leave the column at each of the points **A**, **B** and **C**? [1]
- (b) Explain how the fractionating column separates the crude oil mixture. [3]
- (c) Octane,  $C_8H_{18}$ , is a hydrocarbon in petrol. Hexadecane,  $C_{16}H_{34}$ , is one of the hydrocarbons in ship fuel.
- (i) Show by calculation that hexadecane contains a higher percentage of carbon by mass than octane.

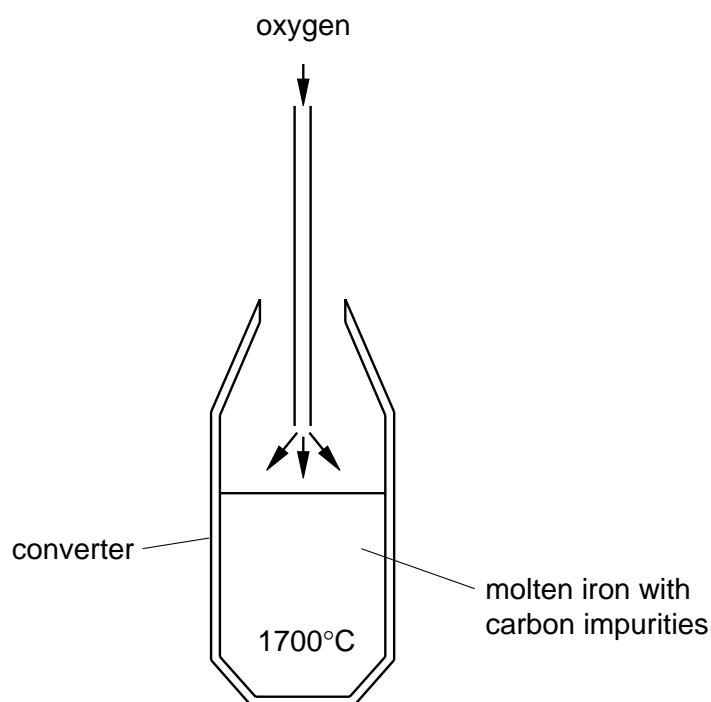
This is the equation for the complete combustion of octane.



- (ii) Write an equation for the complete combustion of hexadecane.
- (iii) Use the equations to explain why hexadecane burns with a smokier flame than octane. [5]
- (d) Name **two** fuels, suitable for cars, which do not come from crude oil. [1]

[Total: 10 marks]

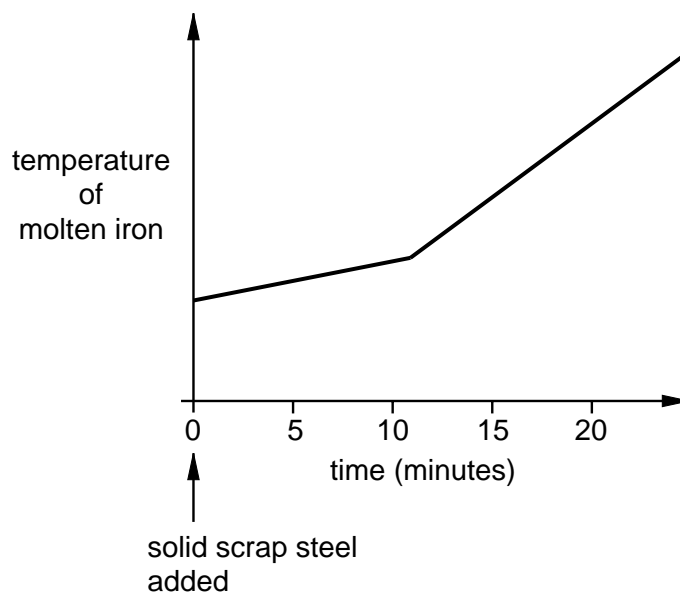
- B9** Iron from the Blast Furnace contains carbon as an impurity. To remove the carbon, oxygen is blown on the molten iron in a large vessel known as a converter. The carbon is oxidised to carbon dioxide.



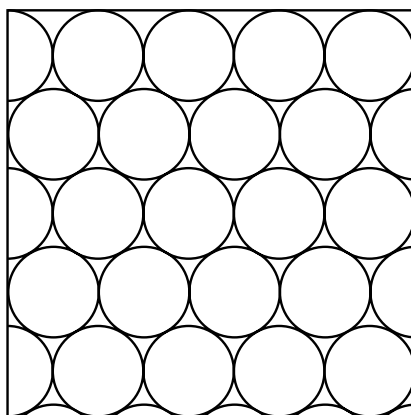
- (a) The temperature of the molten iron increases as the oxygen is blown onto it. Explain why. [1]
- (b) During the oxygen blow, some of the molten iron is oxidised to iron(III) oxide. Write an equation for this reaction. State symbols are not required. [2]

**B9 CONTINUES OVERLEAF.**

- (c) Scrap steel is recycled by being added, as a solid, to the molten iron, before the oxygen blow. The graph below shows how the temperature of the molten iron changes during the oxygen blow.



- (i) Describe how the temperature of the molten iron changes during the oxygen blow. Explain why the solid scrap steel affects the temperature change during the oxygen blow.
- (ii) Give a reason why it is important to recycle steel. [3]
- (d) The diagram shows the arrangement of atoms in pure iron.



Draw similar diagrams to show the arrangement of atoms in

- (i) low carbon steel alloy,
- (ii) high carbon steel alloy.
- (iii) How do the properties of the two types of steel differ? Use your diagrams to explain why the properties are different. [4]

[Total: 10 marks]

**B10** Electroplating can be used to coat nickel with a thin coating of silver.

- (a) Draw a labelled diagram of an apparatus that can be used to electroplate silver onto nickel. [2]
- (b) Write equations, with state symbols, for the reactions at the anode and cathode. [2]
- (c) Solutions of two salts, **A** and **B**, were electrolysed using carbon electrodes. The following products were collected.

<i>salt</i>	<i>products</i>
<b>A</b>	oxygen and hydrogen
<b>B</b>	chlorine and hydrogen

- (i) Suggest the names of the two salts, **A** and **B**. [5]
- (ii) Describe tests to confirm the identifies of the three gases collected. [5]

[Total: 10 marks]

A series of horizontal dotted lines for writing, spanning the width of the page.

A series of horizontal dotted lines for writing, spanning the width of the page.

